

①

ALKENE

① General Formula — C_nH_{2n}
Where n = Number of Carbon atom

- ② Carbon Carbon double bond.
- ③ Sigma bond and pi bond present.
- ④ Hybridisation is sp^2 .
- ⑤ Bond angle is 120° .

pi bond : —

pi bond can be formed by lateral overlapping of orbital.



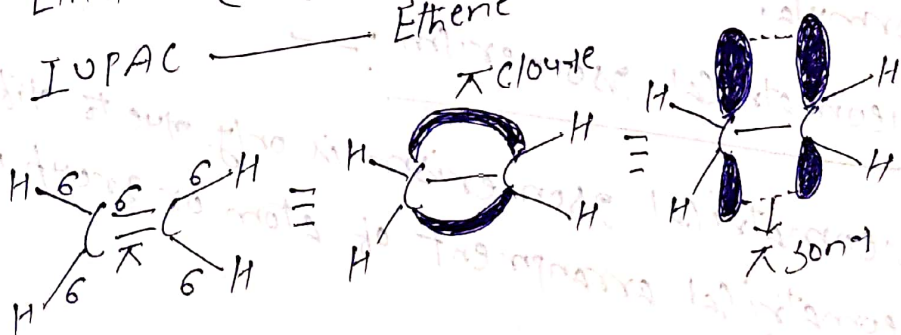
π bond formation

$n=1 \times$

$n=2$

Ex Ethelene (C_2H_4) — Common name.

IUPAC — Ethene



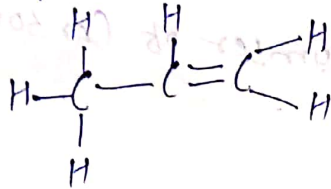
No. of σ bond = 5

No. of π bond = 1

(2)

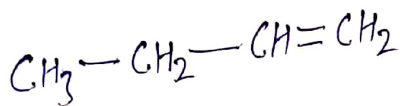
$n=1$

propene (C_3H_6)



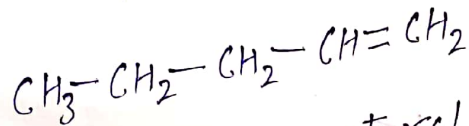
$n=4$

Butene (C_4H_8)



$n=5$

pentene (C_5H_{10})



* Alkene show structural isomerism and Geometrical isomerism.

Structural isomerism:—

Structural isomerism is due to different structural formula.

Geometrical isomerism:—

Geometrical isomerism is only due to different geometrical arrangement of atom or group in space.